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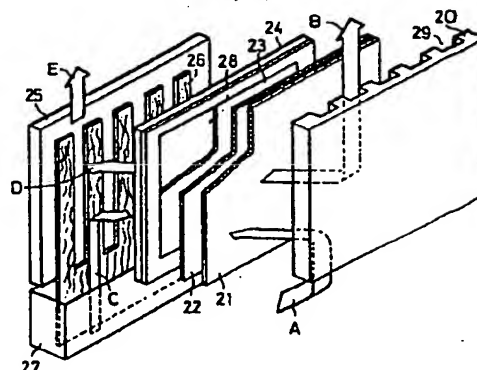
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64 Fuel cell using organic high-molecular electrolyte.

67 In a fuel cell using as an electrolyte an organic, high-molecular electrolyte (23), leakage of the electrolyte (23) from the electrolyte chamber is prevented so that an electrical short-circuit between anode (28) and cathode (21) can be avoided.

FIG. 2



FUEL CELL USING ORGANIC HIGH-MOLECULAR ELECTROLYTE
Field of the Invention

The present invention relates to a novel fuel cell using an organic, high-molecular electrolyte and, more particularly, to a fuel cell which contains a novel electrolyte and which operates at room
5 temperature or at a temperature of at most about 100°C.

Moreover, the present invention relates to an electrolyte structure and an electrolyte composition which are suitable for the above fuel
10 cell.

The present invention is suitable for application both to fuel cells which use a gaseous fuel such as hydrogen or those which use a liquid fuel such as methanol or hydrazine, and for gaseous
15 oxidizers such as oxygen or air or liquid oxidizers such as hydrogen peroxide. However, the present invention should not be limited to such fuel cells.

Description of the Prior Art

A fuel cell is designed to extract the energy
20 generated by the electrochemical reaction of a fuel and an oxidizer directly as electric energy, and is either being researched or is already

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practised as either an electric power plant or a power source for aerospace devices, unmanned facilities on sea or shore, stationary or movable radios, automobiles or domestic appliances.

5 Fuel cells can be divided roughly into representative types such as a molten salt electrolyte types of fuel cell which operate at a high temperature (about 500 to 700°C), phosphoric electrolyte types of fuel cell which operate in the vicinity
10 of 200°C, or alkaline or acidic electrolyte types of fuel cell which operate at room temperature or at a temperature of at most about 100°C.

For the electrolyte of high-temperature fuel cells, carbonates such as K_2CO_3 , which is solid at
15 room temperatures, have frequently been used.

Phosphoric fuel cells use phosphoric acid having a melting point of about 42°C, or polyphosphoric acid, as disclosed in the specification of Japanese Patent Laid-Open No. 5286/1982, both
20 of which are solid or semi-solid at room temperatures.

The electrolyte frequently used in alkaline or acidic electrolyte types of fuel cell used at

a temperature of at most 100°C is either an aqueous solution of caustic potash, caustic soda or lithium hydroxide, or dilute sulfuric acid. This is because aqueous solutions of these strong electrolytes are the most convenient for use because they have a high ionic conductivity at a low temperature.

Those strong electrolytic substances are so corrosive that they restrict the materials that can be used for composing the cell. This means that sufficient care must be taken to prevent those electrolytes from leaking from a cell. However, counter-measures against leakage of electrolyte in existing cells are not very simple so that a variety of trials have been made in the prior art.

The basic problem is raised from the fact that the electrolyte is liquid when the fuel or oxidizer is liquid. The phenomenon occurs that the electrolyte which should stay in the elec-

trolyte chamber passes through a porous fuel or oxidizer electrode into the fuel or oxidizer chamber as a result of dilution due to the concentration gradient within the liquid fuel or
5 oxidizer.

In a fuel cell using a liquid fuel, it is customary as a counter-measure to supply the fuel chamber with a fuel mixture (which is usually called an anodic electrolyte) which has been diluted with
10 the electrolyte. The difference in the concentration of the electrolyte is so reduced by this that outflow of the electrolyte from the electrolyte chamber to the fuel chamber is reduced. Nevertheless, the dilution of the fuel with the electrolyte
15 is a counter-measure which is unnecessary for the intrinsic functions of the cell, the concentration of the fuel is also reduced accordingly so that more power is consumed for circulating the electrolyte than the fuel, with a resultant reduction in the energy efficiency. Moreover, the
20 supply or circulation of the highly corrosive

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electrolyte together with the fuel is disadvantageous for the user in addition to the limitations on the constructional materials.

Although there is an example in which an
5 inorganic powder is mixed with the electrolyte to make a paste, this method has not succeeded in providing a basic counter-measure because the electrolyte will essentially allow the fuel or oxidizer electrode to pass therethrough.

10 Summary of the Invention

It is an object of the present invention to provide a fuel cell which can prevent or reduce leakage or outflow of the electrolyte thereof.

Another object of the present invention is
15 to provide a novel fuel cell which does not use a highly corrosive electrolyte.

A third object of the present invention is to provide an electrolyte structure and an electrolyte composition which are suitable for the
20 above fuel cell.

The present invention is based upon the concept of the principle that the use of an

electrolyte which is incapable of passing through a porous fuel or oxidizer electrode makes it possible to prevent outflow or leakage of the electrolyte in a liquid fuel or oxidizer type of fuel cell of the prior art. The inventors have examined the use of an organic, high-molecular electrolyte as one having the above actions and have confirmed better effects than those anticipated after the assembly of a methanol fuel cell and power-generation experiments.

Description of the Invention

a. Organic, High-Molecular Electrolyte

An organic, high-molecular electrolyte is a well-known substance that is disclosed in detail in "High-Molecular Electrolytes" edited by the Editing Committee of High-Molecular Experiments of the High-Molecular Association and published in 1978 by Kyoritsu Shuppan K.K.

A high-molecular electrolyte dissociates in an aqueous solution to produce large high-molecular ions with high charges and low-molecular ions. The high-molecular electrolyte itself in the solution becomes high-molecular ions (i.e., polyions) which have opposite charges to those of the low-molecular ions, while liberating the low-molecular

ions from their dissociated radicals. The low-molecular ions thus liberated are called "counter-ions". The high-molecular ions have charges at levels corresponding to the number of dissociated
5 ionized radicals.

The counterions thus liberated through dissociation from the high-molecular ions are trapped around the high molecules by the action of an electric field resulting from the high charges
10 of the high-molecular ions of the opposite polarity. If the counterions are trapped or restrained around the high-molecular ions merely by electrostatic action, the counterions thus trapped can migrate freely within the trapped layer. This restraint
15 is called "ø bonding".

On the other hand, the bonding by which the counterions are trapped by the high molecules in the form containing covalent bonds with the dissociated radicals is called "P-bonding". In this
20 case, the counterions are restrained by specific dissociated radicals so that they cannot migrate freely.

Since the high-molecular ions have large molecules, it is thought that their migrations in
25 the solution of the high-molecular electrolyte are

restricted so that they cannot migrate from the electrolyte chamber either to the outside or into the fuel or oxidizer chamber.

As has been described above, by making use
5 of both the limited restraint on the counterions and the restriction on the migration of the high-molecular ions, the fuel cell according to the present invention can provide a number of useful effects, as will be described later in more detail.

10 The kinds and the methods of regulating the high-molecular electrolyte are disclosed in detail in Section 2 of "High-Molecular Electrolytes". In this publication, high-molecular electrolytes are classified into those which are produced from
15 natural high molecules, those which are produced from synthetic high molecules, and special ones (all these are synthetic, high-molecular electrolytes).

The organic, high-molecular electrolytes
20 suitable for use in the present invention are limited to those which can produce as many counterions as possible, namely, those having high electron conductivities. High-molecular acids and bases of strong electrolytes are especially suitable
25 for the present invention.

Suitable as electrolytes are both high-molecular bases, the dissociated radicals of which have basic groups such as amino groups or quaternary ammonium groups such as $-\text{CH}_2\text{N}(\text{CH}_3)_3\text{Cl}$ or
5 $-\text{CH}_2\text{N}(\text{CH}_2\text{CH}_2\text{OH})(\text{CH}_3)_2\text{Cl}$, and high-molecular acids which have acidic groups such as sulfonic groups or phosphine groups.

The residue of the high-molecular electrolyte may be polystyrene, polyethylene, polyacryl,
10 polymethacryl, a styrene-butadiene copolymer or a styrene-divinylbenzene copolymer, which is either sulfonated with sulfuric acid or is subjected to a chloromethylating treatment with a tertiary amine into a quaternary ammonium group.

15 This high-molecular electrolyte has no particular molar weight, but it is set to use a high-molecular substance having a molar weight of at least 1,000, particularly 10,000, in order to prevent its outflow or leakage from the elec-
20 trolyte chamber. In this respect, an ordinary synthetic high-molecular substance has a molar weight of at least 100,000 so that no practical problem is raised.

b. Electrolyte Composition

25 The high-molecular electrolyte is partially

or completely dissolved in water, but when it fills the electrolyte chamber of the fuel cell, it is dissolved either by itself or partially or completely in water in accordance with either
5 necessity or the type of fuel cell.

When the high-molecular electrolyte is used by itself, it may take the form of a powder, fine particles or pellets or any combination thereof. For a methanol fuel cell, for example, the high-
10 molecular electrolyte will swell into or be dissolved in the water before long because methanol and water are mixed and supplied to the fuel.

If a thickening agent is mixed in to increase the viscosity of the solution or paste of a high-
15 molecular electrolyte obtained by swelling, dissolving or dispersing the electrolyte in water, it becomes feasible to apply the solution or paste onto an electrolyte holding frame. For the thickening agent, used either a fine inorganic powder
20 having a high specific resistance or an insulating property, for example a fine powder of silicon carbide, or an organic powder can be used. The term "thickening agent" is made merely for convenience, silicon carbide powder plays the role
25 of a spacer for holding the electrolyte layer and

the fuel or oxidizer electrode at a constant spacing. Hence, the thickening agent could also be called a "spacer material". It is best if this thickening agent is inactive to the electrolyte
5 and water.

If another, low-molecular, electrolyte is added to the solution or paste of the high-molecular electrolyte, it is dissociated to produce positive and negative ions. If only a small quantity thereof
10 is added, the restraint of the above counterions is weakened without losing the features of the high-molecular electrolyte so that the counterions become able to migrate.

Further additions of a variety of additives,
15 if necessary, will not depart from the spirit and scope of the present invention.

c. Electrolyte Structure

The ordinary fuel cell is constructed by...
preparing single cells of a combination of a fuel
20 chamber - a fuel electrode - a electrolyte layer -
an oxidizer electrode - an oxidizer chamber and
connecting a plurality of this single cells in
series to obtain the desired voltage. If a single
cell has an electromotive force of 0.6 volts, a
25 fuel cell having an electromotive force of 12

volts can be constructed by connecting twenty single elements in series. Therefore, each of the components making up a single cell should be shaped as a plate having as small a thickness as possible.

In one example of the present invention, the electrolyte structure can be made thinner if the, above electrolyte composition is carried in a dry or paste form on a frame having a size of 0.1 to 5 mm, preferably 0.3 to 2 mm. It is also effective to apply an organic, high-molecular electrolyte composition to the adjacent sides of the oxidizer electrode and/or the fuel electrode.

In a methanol fuel cell, methanol will pass through the fuel electrode and leak into the electrolyte chamber until it reaches the oxidizer electrode where it will oxidize or burn. In order to prevent this, it is desirable to interpose a diaphragm restraining the passage of the methanol between the fuel electrode and the electrolyte chamber. An ion exchange film, for example, can be used as the diaphragm.

In one embodiment of the present invention, the thickness of the electrolyte structure is greatly reduced, and the assembly of the cell is

facilitated, by attaching an ion exchange film to **0100530**
the electrolyte holding frame and by carrying the
above electrolyte composition on either or both
sides of the ion exchange film.

5 The electrolyte holding frame may be suitably
made of an insulator which may be any of a variety
of plastic plates, sheets or films, for example.
It is possible to prevent the electrodes from
short-circuiting if the above electrolyte composi-
10 tion mixed with a thickening agent or spacer mate-
rial is used.

BRIEF DESCRIPTION OF THE DRAWINGS

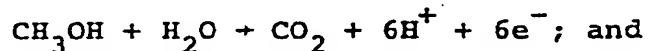
Fig. 1 is a schematic view explaining the
principle of the fuel cell according to the present
15 invention; Fig. 2 is a partially cut-away perspec-
tive view showing the construction of a single
cell of the fuel cell according to the present
invention; Fig. 3 is a perspective view showing
the structure of a fuel battery which is constructed
20 by connecting single cells in series; Figs. 4 to
6 are graphs illustrating the V - I characteristics
of a single cell of the fuel battery according to
embodiments of the present invention; Fig. 7 is
a partially cut-away perspective view showing
25 the construction of an electrolyte structure
according to the present invention; and Fig. 8 is
a sectioned perspective view explaining the fuel
supply and control method of an embodiment of the
present invention.

DETAILED DESCRIPTION OF PREFERRED EMBODIMENTS

Fig. 1 is a schematic view showing the principle of the fuel cell according to the present invention, in which a fuel chamber 6 is supplied with a fuel 5 and an oxidizer chamber 8 is supplied with an oxidizer 7 such as oxygen, air or hydrogen peroxide. An organic, high-molecular electrolyte layer 4 is formed between a fuel electrode 2 and an oxidizer electrode 3. The high-molecular electrolyte dissociates in the presence of water into high-molecular ions 15 and ionized radicals 14. These ionized radicals 14 migrate within the electrolyte in accordance with the cell reactions to receive the electrons generated at the fuel electrode, and carry them to the oxidizer electrode. The high-molecular ions migrate neither to the oxidizer electrode nor to the fuel electrode because they have large molecular weights.

If a methanol fuel cell is taken as an example, the following cell reactions occur:

at the methanol electrode (i.e., the ANODE):



at the oxidizer electrode (i.e., the CATHODE):

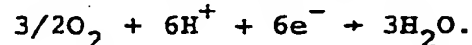


Fig. 2 is a perspective view showing the construction of a single cell of a methanol-air fuel battery according to one embodiment of the present invention.

As shown, a single cell is constructed by **0100530**
consecutively juxtaposing a separator 20 of graphite
which forms an air chamber and acts as a collector,
an air electrode 21 adjacent to the separator 20, ion
5 exchange membrane 22, a layer 23 of a paste of
a high-molecular electrolyte composition carried
on a methanol electrode 28, and a separator 25 of
graphite which forms a fuel chamber and acts as a
collector. The separator 20 is formed with grooves
10 29 to provide air passages.

In the embodiment shown, a fibrous wick member
26 is provided which is used to supply the methanol
from a methanol tank 27 to the fuel chamber by
capillary action. This embodiment is featured
15 by eliminating auxiliary machinery such as a pump
for supplying or circulating the anolyte, so is
different from the methanol fuel cell of the
prior art. If a pump is used, power for driving
the pump is required not only during the start but
20 also during running. The pump can be dispensed
with if the cell has the construction shown in
Fig. 2.

The reason for the fuel supply method shown
in Fig. 2 is that it makes it unnecessary to use
25 an anolyte (i.e., the mixture of fuel and elec-
trolyte). In the dilute sulfuric acid electrolyte
type of methanol cell of the prior art, a large
amount of sulfuric acid (20 to 70 by volume) has
to be mixed with the anolyte. As a result, the

wick system shown in Fig. 2 is difficult to adopt for supplying the necessary methanol because the concentration of the methanol is very low.

In the present invention, however, an anolyte
5 need not be used, either methanol alone, or methanol to which a small quantity of water necessary for the reactions is added, can be used so that sufficient methanol can be supplied even by a wick system. As a result, auxiliary machinery for
10 the fuel supply can be dispensed with so that the efficiency of the energy use can be increased.

In the figure, incidentally, arrow A indicates the flow of air, arrow B the flows of steam and air which are generated by the cell reactions,
15 arrow C the flow of fuel, arrow D the flow of carbonic acid gas which is generated by the cell reactions, and arrow E the flow of carbonic acid gas which is discharged from the cell.

The ion exchange film 22 may be sandwiched
20 between the methanol electrode 28 and the electrolyte paste 23. The methanol electrode 28 is held on a frame 24.

Fig. 3 is a perspective view showing the appearance of the fuel battery which is constructed
25 by laminating the single cells of Fig. 2. In Fig. 3, the same reference characters as those of Fig. 2 indicate the same parts and portions. The plurality of cells thus laminated are accommodated

in a cell case 31 which has positive and negative terminals 32 and 33 attached thereto. Since the air supply system of Fig. 2 depends on natural convection, an epoch-making fuel cell is provided in that there is no requirement for auxiliary machinery supplying the fuel and air. The fuel is supplied from a port 30 to the tank 27.

Figs. 4, 5 and 6 are graphs illustrating the V - I characteristics of single cells according to various embodiments of the present invention.

Fig. 7 shows the construction of the electrolyte structure according to another embodiment of the present invention. The embodiment shown in Fig. 2 is constructed so that an electrolyte composition on which an ion exchange film is arranged is applied to the surface of the fuel electrode. In Fig. 7, on the other hand, the ion exchange film or diaphragm 43 is sandwiched between electrolyte frames 41 and 42 to be held thereby, and has a high-molecular electrolyte composition 44 applied to both or either of its sides. This construction makes it feasible to assemble a single cell.

Fig. 8 is a sectioned perspective view showing one example of a method of supplying the fuel. Either fuel or a mixture of fuel and water in a fuel tank 57 is supplied by capillary action to a fuel electrode 52 by means of a transparent fibrous wick 56. An electrolyte layer 51 is formed between

the fuel electrode 52 and an oxidizer electrode 53. On the outer side of the oxidizer electrode 53, a conductive spacer 55 formed with oxidizer passages 61 is mounted. As shown in Fig. 8, a stopper 59 is provided which can stop the fuel supply if it is pushed in the direction of the arrow when the operation of the fuel cell is to be interrupted.

Examples

10 Polystyrene sulfonic acid was used as the electrolyte of the fuel cell shown in Fig. 2 which used methanol directly as the fuel thereof.

The polystyrene sulfonic acid was prepared by sulfonation by causing dense sulfuric acid at 100°C to act upon a film of polystyrene in the presence of an catalyst of 0.2% silver sulfate. The polystyrene dissolved into a viscous liquid. After about eight hours, the polystyrene was dissolved in water so that it could be purified to prepare the polystyrene sulfonic acid.

For the sulfonating process, incidentally, a process using a complex or adduct of SO_3 (e.g., $\text{C}_6\text{H}_5\text{COOSO}_3\text{H}$), or Eisenberg's method using 100% H_2SO_4 (as disclosed by W. R Carroll, H. Eisenberg, J. Polymer Sci. 4, 599 (1966) could be used.

An electrolyte paste obtained by kneading the polystyrene sulfonic acid with a suitable moisture content, was sandwiched to a thickness of about 0.2 mm between the two electrodes 24 and 21 which had been treated by a catalyst. Because of the paste form, care was taken to prevent bubbles forming in the interfaces between the electrolyte and the electrodes either by applying the paste to the surfaces of the electrodes or by compressing them after they have been sandwiched together. The current (I) - voltage (V) characteristics of this cell are illustrated in Fig. 4.

In this way, the electrolyte is fixed. As a result, fuel without electrolyte can be supplied so that the outflow of acid is eliminated to facilitate handling. An important feature obtainable because only fuel is supplied to the laminated cells shown in Fig. 3 is that not only the electrodes can be prevented from short-circuiting but also the fuel supply structure can be simplified. Moreover, the fuel can be supplied directly from the fuel tank 27 by simplifying or omitting the fuel supply pump. This can also be applied to fuel cells (with either acidic or basic electrolytes) which have previously used a liquid

electrolyte.

In another example, it is possible to use as the electrolyte of the fuel cell a paste electrolyte which has been kneaded with an electrolyte exhibiting the same acidic nature as that of the high-molecular acid exhibiting the acidic electrolyte of the above fuel cell using methanol as its fuel. For example, polystyrene sulfonic acid and 3 mols of dilute sulfuric acid were kneaded into a paste which was sandwiched between the two electrodes 21 and 24 which had been treated with a catalyst. The thickness after this treatment was 0.2 mm. Inorganic particles such as silicon carbide can be treated with a suspension of a fluorine resin and kneaded with the above polystyrene sulfonic acid and 3 mols of dilute sulfuric acid into paste. In this case, the suspension of fluorine resin acts as a binder and forms an electrolyte reserving layer so that it can act as an excellent electrolyte. This results in the excellent cell characteristics shown in Fig. 5.

In still another example, a paste of a high-molecular acid (such as polystyrene sulfonic acid) is kneaded with a bead-shaped, high-molecular acid

(such as a styrene-divinylbenzene copolymer sulfonic acid), and the resultant mixture can be used as the electrolyte of the fuel cell. For example, a substance obtained by kneading the
5 above polystyrene sulfonic acid and a suitable moisture content is kneaded with beads of the styrene-divinylbenzene copolymer sulfonic acid (having a particle diameter of a few μm) to prepare a paste, and this paste is used as the
10 electrolyte of the fuel cell. In this case, beads of polystyrene sulfonic acid can be used instead of the above silicon carbide so that the concentration of hydrogen ions in the electrolyte is increased and the beads perform as the reserve layer
15 for the electrolyte. As a result, the internal impedance of the fuel cell is reduced so that the excellent characteristics shown in Fig. 6 are exhibited.

The description thus far made is directed to
20 examples in which a paste of a high-molecular acid is used as the electrolyte of a fuel cell. Nevertheless, such high-molecular acids can also be used in the form of a crosslinked gel.

For example, chlorosulfonic acid can react
25 with crosslinked polystyrene in a solution of

sulfuric acid so that it becomes a gel. By using this gel as the electrolyte of a fuel cell, it is possible to provide an electrolyte which is stable even for use over a prolonged period of time.

5 A combination of a high-molecular electrolyte and an ion exchange membrane can also be used as an electrolyte. For example, the high-molecular acid of polystyrene sulfonic acid and a film or a cation exchange film containing both styrene-
10 butadiene or styrene divinylbenzene in its parent body and a sulfonic group are combined with each other to prepare the electrolyte for a fuel cell. In this case, the advantage can be obtained that the ion exchange film prevents the fuel from
15 migrating from the fuel electrode 24 to the air electrode 21 so that the efficiency of fuel usage can be improved.

 The above description is directed mainly to examples of fuel cells in which a liquid (i.e.,
20 methanol) using an acidic electrolyte is used as the fuel. In this case, however, polyethylene sulfonic acid or the like can be used by a similar process in addition to the use of polystyrene sulfonic acid as the electrolyte of the fuel cell.

25 A high-molecular acid can be used in a fuel

cell which employs a gas (such as hydrogen) as the fuel with the same acidic electrolyte.

A high-molecular electrolyte can be used as the electrolyte of a fuel cell using a basic
5 electrolyte.

At present, hydrogen, hydradine, alcohol or the like is used as fuel whereas caustic potash is used as an electrolyte. A basic electrolyte such as polyvinylpyridine can be made into a
10 paster or gel in the same way as in the above description so that it can be used as the electrolyte of a fuel cell.

A basic, high-molecular electrolyte such as polyvinylbenzyltrimethyl ammonium may be used.
15 This electrolyte produces a quaternary ammonium salt if the polystyrene is chloromethylated and treated with a tertiary amine.

First 1% of polystyrene and 0.3% of $ZnCl_2$ acting as a catalyst are dissolved in purified
20 chloromethyl ether. This solution is chloromethylated at 50°C for 2 hours, and the resultant product is purified and dried.

The chloromethylated polystyrene is dissolved in dimethylformamine, and twice the equivalent
25 amount of trimethylamine is blown into the resultant

solution so that the reaction is conducted at room temperature (e.g., 15°C) for one day. The product is poured into ether so that it is deposited and purified.

5 Fig. 8 shows one embodiment (1) of the fuel supply method according to the present invention. The fuel electrode 52 and the oxidizer electrode 53 face each other across the electrolyte chamber 51, the fuel chamber 54 is arranged on the side
10 of the fuel electrode 52 and the oxidizer chamber 55 is arranged on the side of the oxidizer electrode 53. In the fuel chamber 54, the fuel wick 56, which is made of a fibrous material of matted, woven or unwoven form such as carbon, acryl,
15 polyamide or polypropyrene, extends longitudinally from the fuel electrode 52 so that its lower end face is in contact with the bottom of the fuel tank 57.

 When the fuel tank 57 is filled with liquid
20 fuel 58, this fuel 58 rises through the fuel wick 56 by capillary action to fill the fuel chamber 54 with the liquid fuel 58 and supply the liquid fuel 58 to the fuel electrode 52 through the sections of the fibrous material 56, so that
25 electric power is generated by the electrochemical

reaction. According to the present embodiment,
an effect can be obtained that the liquid fuel 58
to be consumed is supplied continuously from the
fuel tank 57 through the fuel wick 56 so that
5 power generation can continue while there is
liquid fuel 58 in the fuel tank 57.

The thickness of the fibrous member 56 is
determined either by the suction characteristics
of the liquid fuel 58 by the fibrous member 56,
10 for example, the shape, size, etc., of the fuel
electrode, or by the energy generated by the cell,
i.e., the consumption of the fuel 58, but may
desirably be between about 10 μ m to 10 mm from a
practical standpoint.

15 One application of a fuel supply making use
of capillary action is the wick of an oilstove.
In this case, however, the fuel sucked up from
the fuel tank is gasified and burned at the tip
of the wick. In the present invention, however,
20 the fuel introduced from the fuel tank is supplied
to the electrode surfaces from the surface of a
material having capillary tubes. Thus, the present
invention has an effect different from that of the
wick of an oilstove.

25 The fuel chamber 54 can be made integral with

the fuel tank 57.. The space between the fuel chamber 54 and the fuel tank 57 is filled either with the fibrous wick 56 of matted, woven or unwoven carbon, acryl, polyamide or polypropylene, or a porous, inorganic fuel wick, which has been sintered of an inorganic powder of alumina or silica as to have a number of pores, such that it extends at least from the lower end of the fuel electrode 52 to the bottom of the fuel tank 57.

When the fuel tank 57 is filled with the liquid fuel 58, this liquid fuel 58 is sucked from the fuel tank 57 through the fibrous member 56 or the porous, inorganic fuel wick by capillary action so that it is diffused into the fuel chamber by osmotic pressure until it is supplied to the fuel electrode 52 for powder-generating purposes.

According to this embodiment, the liquid fuel is diffused at the rate of consumption by the fuel electrode 52 from the fuel tank 57 through the wick 56 or an inorganic fuel wick 56 into the fuel chamber 54 so that power generation continues while there is fuel liquid 58 in the fuel tank 57.

In a different embodiment, the fuel electrode 52 is constructed of an integrated structure of

the porous substrate and the catalyst layer, and a porous substrate is used as the fuel wick. By connecting the wick to the fuel 58 in the fuel tank 57, it is possible to supply the fuel electrode 52 with the fuel. This results in the advantages that the construction is simplified and the products can be discharged easily from the fuel electrode.

The fuel cell of Fig. 8 is equipped with a mechanism for controlling the rate of flow of the fuel supplied. A fuel controlling fastening 59 may also be provided for applying a mechanical pressure in the direction of the arrow to the fuel wick 56 rising from the fuel tank 57 which is interposed between the fuel tank 57 and the fuel chamber 54, thereby suppressing the supply of the fuel 58.

By compressing the fuel wick 56 by the use of the fuel controlling fastening 59, it is possible to reduce or stop the fuel being supplied from the fuel tank 57 to the fuel chamber 54 by capillary action.

When the power generation is to be interrupted, therefore, by the use of this mechanism, the handling can be simplified because the fuel 58 can be

left in the fuel tank 57.

The fuel supply can also be conducted more effectively in the following manner.

The fuel tank 57 is enlarged in the direction
5 of the height thereof parallel to the cell structural parts such as the fuel electrode 52 or the oxidizer electrode 53, and is filled with the fuel 58. In this case, a pressure corresponding to the level of the liquid 58 is applied to the
10 fuel wick 56 so that the height capable of supplying fuel is increased accordingly. As a result, even if the height of the battery is large when the area of the cell required is large, fuel supply is facilitated to provide the effect that a drop
15 of the output due to a shortage of fuel can be prevented.

The above description of the embodiments has laid stress upon a methanol-air fuel cell. However, the present invention should not be limited
20 to those embodiments but can be applied to different fuel cells using other liquid fuels and oxidizers.

Similarly, the above description of the electrolyte composition has concentrated on a paste of an organic, high-molecular electrolyte composition.
25 However, the fundamental object of the present

invention to prevent leakage or outflow of the electrolyte can be achieved even if the organic, high-molecular electrolyte is in a solution.

Since the outflow of the electrolyte from
5 the electrolyte chamber is prevented according to the present invention, a prolonged stability of the cell characteristics can be obtained, and in addition the effect of preventing the electrolyte from leaking to the outside of the cell can be
10 further improved. As a result, a suitable effect can be obtained especially for moving power sources and batteries which must be used in places subject to vibrations. Moreover, handling as a domestic power source is facilitated to provide a suitable
15 power source.

With a sulfuric acid electrolyte or an alkali electrolyte, furthermore, the mixed solution (i.e., the anolyte) of the liquid fuel and the electrolyte can be supplied without any pump, differentiat-
20 ing it from the prior art in which the mixed solution is supplied by means of a pump. Since the pump capacity is usually about one mol of the fuel concentration, the energy consumed by the pump is much higher than that required for the supply of
25 fuel alone. In this respect, according to the

present invention, the supply of fuel alone or fuel with a small amount of water, is made possible so that auxiliary machinery can be reduced. Furthermore, power losses due to short-circuiting
5 between the electrodes by the electrolyte is avoided in to the present invention to provide a large effect in view of the improvement in the efficiency and in the simplification of the construction. When the fuel is gaseous, similarly,
10 the outflow of liquid cannot be avoided if this electrolyte is used. According to the present invention, however, the outflow of electrolyte is eliminated so that the life of the cell can be extended.

CLAIMS:

1. A fuel cell of the type wherein an electrochemical reaction is caused between a fuel electrode (28) and an oxidizer electrode (21) facing each other via an electrolyte (23) so that electric energy may be extracted
5 directly, characterised in that said electrolyte (23) is an organic high-molecular electrolyte that will dissociate in the presence of water.
2. A fuel cell as set forth in claim 1, wherein at least a part of a fuel or an oxidizer is liquid when said fuel
10 cell is operating.
3. A fuel cell as set forth in claim 1, wherein said fuel is liquid and contains water.
4. A fuel cell as set forth in claim 1, wherein said high-molecular electrolyte (23) is of a strong acid or base
15 type.
5. A fuel cell as set forth in claim 1, wherein said high-molecular electrolyte (23) contains water in which it is partially dissolved.

6. A fuel cell as set forth in claim 1, wherein said high-molecular electrolyte (23) is in the form of a paste containing water and a thickening agent.
7. A fuel cell as set forth in claim 1, wherein said high-molecular electrolyte (23) is mixed with an inactive thickening agent.
8. A fuel cell as set forth in claim 7, wherein said inactive thickening agent is an inorganic powder having a high specific resistance.
9. A fuel cell as set forth in claim 1, wherein said high-molecular electrolyte (23) has a dissociated radical of a sulfonic or phosphine group.
10. A fuel cell as set forth in claim 1, wherein said high-molecular electrolyte (23) has a dissociated radical of an amino group, a quaternary ammonium group, $-\text{CH}_2\text{N}(\text{CH}_3)_3^+$ or $-\text{CH}_2\text{N}(\text{CH}_2\text{CH}_2\text{OH})(\text{CH}_3)_2\text{Cl}$.
11. A fuel cell as set forth in claim 1, wherein at least one of the adjacent sides of said fuel electrode (28) and said oxidizer electrode (21) carry said organic, high-molecular electrolyte (23) in the form of a film.

12. A fuel cell as set forth in claim 1, comprising means (26) for utilizing capillary action to supply said fuel electrode (28) with fuel or a mixture of fuel and water.

13. A fuel cell as set forth in claim 1, wherein a high-
5 molecular residue of said high-molecular electrolyte (23) is polystyrene, polyethylene, polyacryl, polymethacryl, a styren= butadiene copolymer or a styrene-divinylbenzene copolymer.

The diagram illustrates the internal structure and carrier transport in a BJT. The emitter (1) is on the left, the base (2) is in the middle, and the collector (3) is on the right. The base is divided into two regions (4 and 5) by a vertical boundary. The emitter and collector are connected to external circuits (6 and 7) which provide a forward bias. Arrows show the movement of electrons (indicated by '-' signs) from the emitter into the base and towards the collector, and holes (indicated by '+' signs) from the collector into the base and towards the emitter. The regions are labeled with numbers 1 through 8, and the carrier concentrations are indicated by '+' and '-' signs.

[illegible]

2/4

FIG. 3

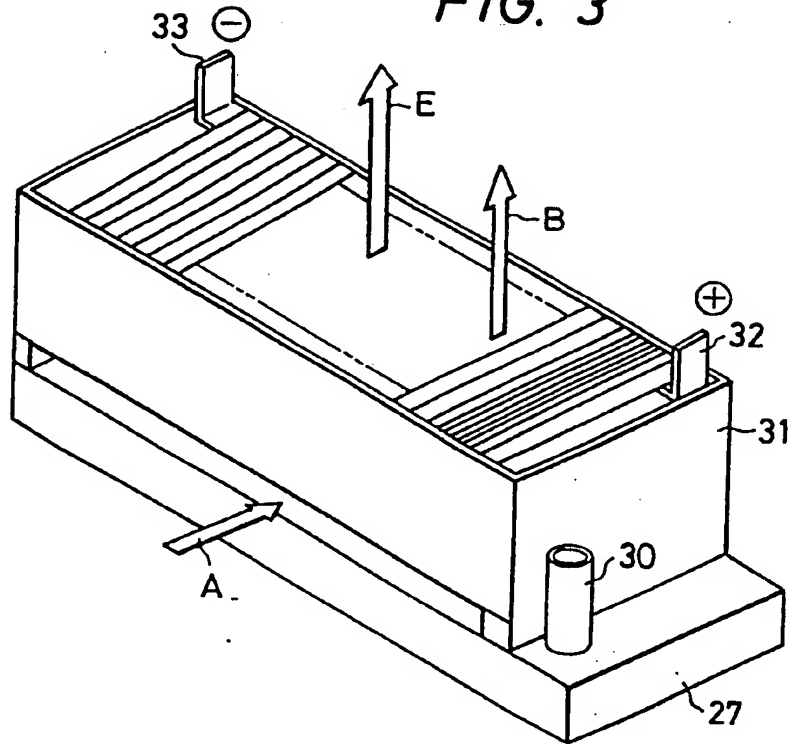
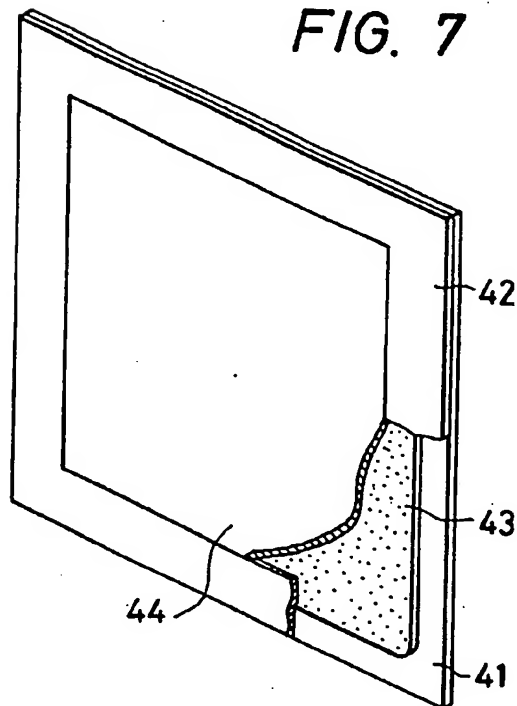


FIG. 7



3/4

FIG. 4

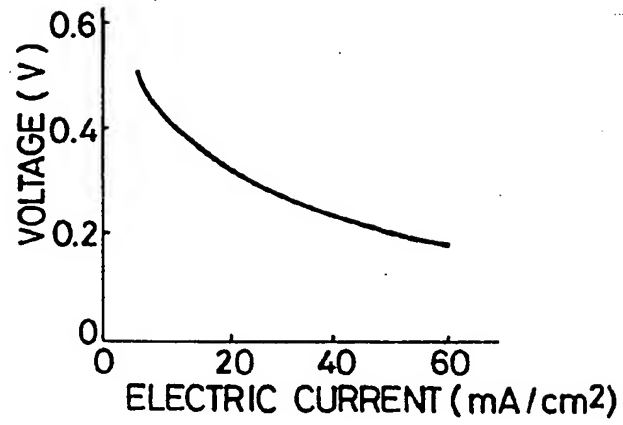


FIG. 5

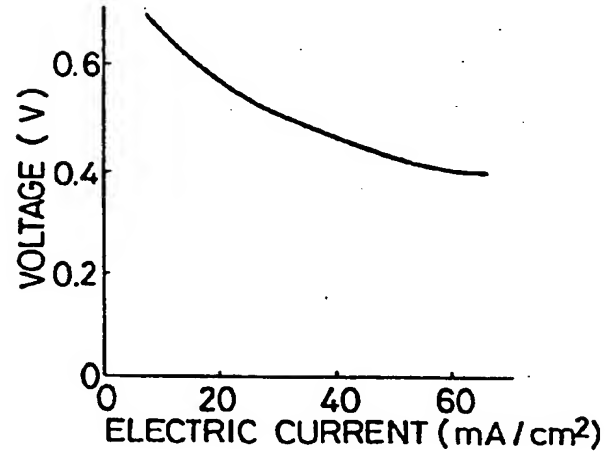
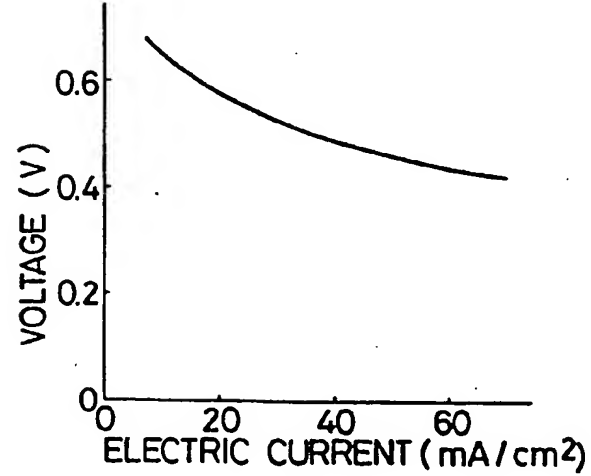


FIG. 6



A/A

FIG. 8

